

2.4: Cornell Note Activities

Tutor Handout 2.4.3 (4 of 4)

Class Notes If there was no class lecture this week, write a paragraph about what you learned and/or questions about what you didn't understand.	Name: <u>Student A</u>
Topic: <u>Compounds</u>	Class: <u>Science</u>
Questions/Main Ideas:	Period: <u>5</u>
	Date: <u>1/12/03</u>
	Notes:
Explain how ions form.	Ions form when atoms become charged after gaining an electron or losing an electron.
How do ionic compounds exist?	Ionic compounds exist in the form of crystals made of many ions, each attracted to the other ions of the opposite charge.
Describe the characteristics of ionic compounds.	Ionic compounds have high melting and boiling points. They conduct electricity well when dissolved in water.
Explain how valence electrons are shared in a covalent bond.	In covalent bonds, valence electrons are shared between the atoms.
When is a polar bond created?	Polar bonds are created when atoms share electrons unevenly.
Which attractions are stronger, polar or nonpolar?	Attractions between polar molecules are stronger than nonpolar attractions.
Summary: Ions form when atoms gain or lose electrons. Ionic compounds resemble crystals and have high melting and boiling points. Polar bonds are attracted more strongly than nonpolar ones.	